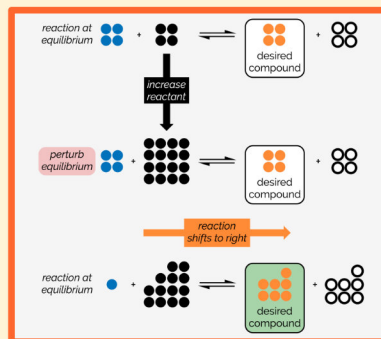


HIGHLIGHTS

- State Le Chatelier's Principle
- Effect of changing concentration
- Effect of changing pressure
- Effect of changing temperature



Reactions are at equilibrium, and, as long as you don't perturb or stress the equilibrium, the concentrations will remain constant. Alter the conditions, and the reaction will shift in one direction or another in order to minimize or counteract the change. This observation is generally known as *Le Chatelier's Principle*. This is a description of the change in reaction not an explanation of the change.

If you understand Le Chatelier's Principle, you can use it to push a reaction in a desired direction and maximize the amount of desired product that you can prepare.

CHEMISTRY CLASSICS

LE CHATELIER'S PRINCIPLE

GETTING WHAT YOU WANT FROM AN EQUILIBRIUM REACTION

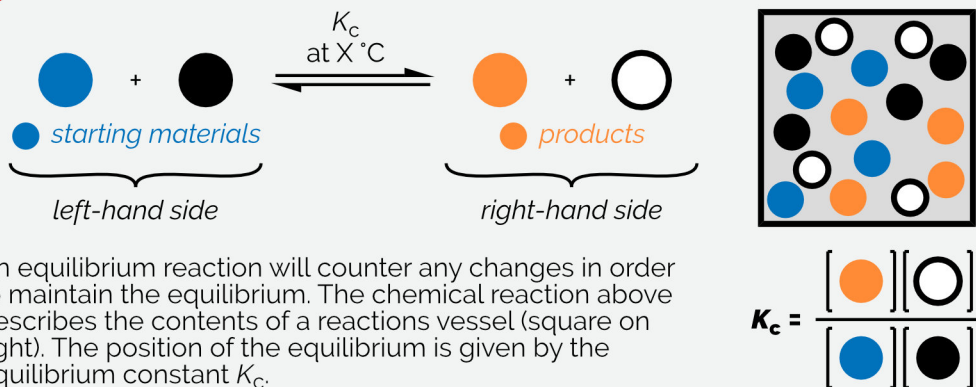


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Le Chatelier's Principle (*getting what you want from equilibrium reactions*)

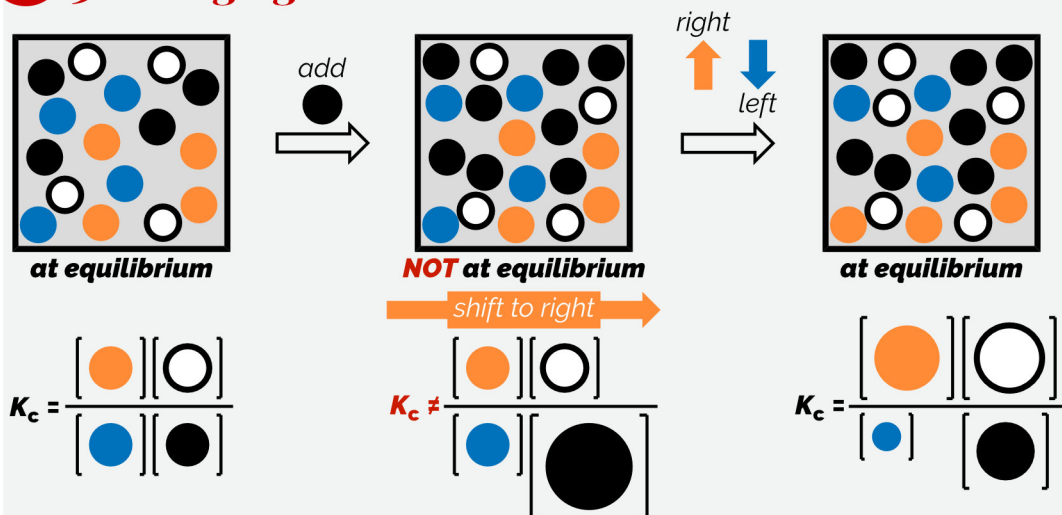
1. General reaction



2. Le Chatelier's principle

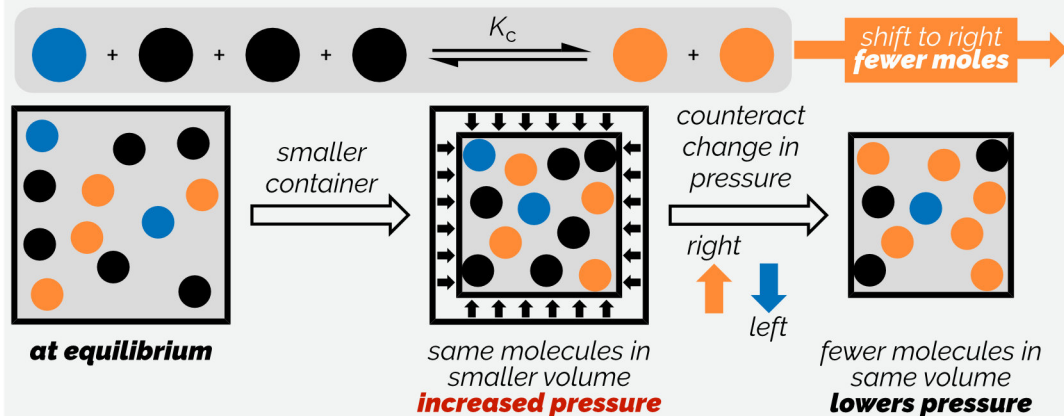
When a chemical system at equilibrium is disturbed, it responds by shifting the equilibrium composition in such a way to counteract the change.

3. Changing concentration of reactants



4. Changing the pressure

At least one reactant must be a gas. There must be different #moles of gas on each side of reaction & pressure is changed by altering volume, not by adding inert gas.



5. Change in temperature

